

(12) INTERNATIONAL APPLICATION PUBLISHED UNDER THE PATENT COOPERATION TREATY (PCT)

(19) World Intellectual Property Organization
International Bureau



(43) International Publication Date
8 November 2001 (08.11.2001)

PCT

(10) International Publication Number
WO 01/83855 A1

(51) International Patent Classification⁷: **C25D 7/00**,
5/10, H01L 39/24

(21) International Application Number: **PCT/GB01/01793**

(22) International Filing Date: 20 April 2001 (20.04.2001)

(25) Filing Language: English

(26) Publication Language: English

(30) Priority Data:
0010494.3 28 April 2000 (28.04.2000) GB

(71) Applicant (for all designated States except US): **ISIS IN-
NOVATION LIMITED** [GB/GB]; Ewert House, Ewert
Place, Summertown, Oxford OX2 7SG (GB).

(72) Inventors; and

(75) Inventors/Applicants (for US only): **MOORE, Joanne,
Camilla** [GB/GB]; Department of Materials, Univer-
sity of Oxford, Parks Road, Oxford OX1 3PH (GB). **GROVENOR, Christopher, Richard, Munro** [GB/GB];
Department of Materials, University of Oxford, Parks
Road, Oxford OX1 3PH (GB). **GOODALL, Russell**
[GB/GB]; 11 Green Meadow Drive, Parc Seymour,
Caldicot, South Wales NP26 3AW (GB). **WHITELEY,
Richard, Melville** [GB/GB]; 115 Newbury Gardens,
Stoneleigh, Epsom, Surrey KT19 0PS (GB).

(74) Agents: **ELLIS-JONES, Patrick, George, Armine et al.**;
J.A. Kemp & CO., Gray's Inn, 14 South Square, London
WC1R 5LX (GB).

(81) Designated States (national): AE, AG, AL, AM, AT, AU,
AZ, BA, BB, BG, BR, BY, BZ, CA, CH, CN, CO, CR, CU,
CZ, DE, DK, DM, DZ, EE, ES, FI, GB, GD, GE, GH, GM,
HR, HU, ID, IL, IN, IS, JP, KE, KG, KP, KR, KZ, LC, LK,
LR, LS, LT, LU, LV, MA, MD, MG, MK, MN, MW, MX,
MZ, NO, NZ, PL, PT, RO, RU, SD, SE, SG, SI, SK, SL,
TJ, TM, TR, TT, TZ, UA, UG, US, UZ, VN, YU, ZA, ZW.

(84) Designated States (regional): ARIPO patent (GH, GM,
KE, LS, MW, MZ, SD, SL, SZ, TZ, UG, ZW), Eurasian
patent (AM, AZ, BY, KG, KZ, MD, RU, TJ, TM), European
patent (AT, BE, CH, CY, DE, DK, ES, FI, FR, GB, GR, IE,
IT, LU, MC, NL, PT, SE, TR), OAPI patent (BF, BJ, CF,
CG, CI, CM, GA, GN, GW, ML, MR, NE, SN, TD, TG).

Published:

- with international search report
- before the expiration of the time limit for amending the
claims and to be republished in the event of receipt of
amendments

For two-letter codes and other abbreviations, refer to the "Guid-
ance Notes on Codes and Abbreviations" appearing at the begin-
ning of each regular issue of the PCT Gazette.



WO 01/83855 A1

(54) Title: TEXTURED METAL ARTICLE

(57) Abstract: A process for producing a metal article coated with a metal layer having a biaxially textured surface, which process comprises electrodepositing the metal layer on a biaxially textured metal substrate such that the surface of the metal layer has the same texture as that of the substrate.

TEXTURED METAL ARTICLE

This invention relates to biaxially textured metal articles.

There has been a considerable amount of research into the production of wires
5 and other articles from high temperature ceramic superconductors. It is now
accepted that very few types of grain boundary between crystals can support the high
currents required by many applications of superconductors. In practice, this means
that it is desirable to make large areas or long lengths of tape where all the
superconducting grains are in a single orientation and so simulate a single crystal.
10 One way of doing this is by making superconductors as single crystals or in
essentially single-crystal form in the form of epitaxial films on single crystal or
biaxially textured substrates. It is commonly found that a biaxial texture is necessary
to obtain high transport critical current densities.

A typical method for producing such a superconducting material is to make a
15 textured substrate from a metal such as nickel or a nickel alloy. The substrate is
textured by rolling and recrystallising. The crystallographic orientation or texture of
the substrate is then used as a template onto which one or usually two or more
barrier/buffer layers of metal and then the superconductor are deposited in order,
retaining the texture of the metal substrate. The barrier layers are for example metals
20 and oxides or just oxides. The buffer layer can be deposited by a number of methods
typically including vacuum deposition methods, e.g. sputtering, evaporation or dip-
coating of solutions or gels. Once the buffer layer(s) have been deposited, the article
is subjected to a heat treatment in order to impart the texture to the buffer layers. The
superconductor, for example YBaCuO, is then deposited on the buffer layer(s).

25 Generally buffer layers such as silver or oxides such as CeO₂ are deposited on
the substrate using vacuum techniques such as sputtering. Typically an additional
buffer layer of another metal such as palladium or platinum is deposited on the
substrate prior to depositing the silver or oxide buffer layer in order to reduce the
lattice mismatch between silver and the substrate.

30 A typical superconductor with a noble metal buffer layer comprises a
typically 50 - 150µm thick nickel (or alloy) substrate coated with 20 to 2000nm of

-2-

sputtered palladium or platinum. This is coated with a typically 100nm to 25µm thick coating of silver or silver oxide. This is optionally coated with a third buffer layer of a material such as CeO₂, MgO or YSZ (yttria-stabilised zirconia). The substrate and its coatings is generally subjected to a heat treatment step during
5 deposition of the layers in order to cause the coatings to develop the texture of the substrate. The superconducting layer is then deposited on the textured silver surface.

A ceramic buffer layer may be used in the production of a superconductor. Examples of ceramic buffer layers are YSZ, MgO, TiN, ZrO₂, CeO₂, LaAlO₃ and SrTiO₃. These are suitably deposited by techniques such as vacuum coating or sol-
10 gel on the metal coated substrate. The superconductor layer is then deposited on the ceramic buffer layer.

Another process typically uses a nickel or nickel alloy substrate coated with a 10 to 100nm thick layer of CeO₂. The buffer layer of CeO₂ is generally covered with a 50 to 1000 nm thick layer of YSZ or Yb₂O₃. A third buffer layer of for example
15 CeO₂, Yb₂O₃ or LaAlO₃ may also be used. A YBCO superconductor layer is then typically deposited on the combination of buffer layers for example by pulsed laser deposition or BaF₂ precursor co-evaporation. Other suitable superconductors include NdBCO, and Tl, Bi1223. This process may also be used with a platinum-palladium substrate.

20 Other suitable buffer layers include Yb₂O₃, Gd₂O₃, NiO, NdAlO₃ and LaAlO₃. These are generally used on a pure nickel substrate between the substrate and a superconducting layer of for example YBCO.

It is necessary to have the coating layers between the metal substrate, which is typically nickel or copper, and the superconducting layer, for example YBaCuO,
25 because nickel and copper and other suitable substrate materials usually react with the superconducting layer. It is therefore necessary to separate the superconducting layer and the substrate by a buffer layer.

An object of the present invention is to provide a new process for producing biaxially textured metal coated metal articles. A further object of the invention is to
30 use the biaxially textured coated metal articles as substrates in the production of superconducting articles. In particular this invention seeks to provide a process for

coating a metal substrate with a silver buffer coating suitable for further coating with a superconducting layer. In the process of the invention a thin layer of metal is provided on a suitable substrate. The metal is electrodeposited on the substrate forming a thin layer that maintains the texture of the substrate. The invention forms
5 a textured bi-layer or laminate structure where each layer is distinct i.e. there is substantially no diffusion of atoms from one layer into the other layer. Thus the crystallographic texture of the deposited layer follows faithfully that of the substrate.

Accordingly, the present invention provides a process for producing a metal
10 article coated with a metal layer having a biaxially textured surface, which process comprises electrodepositing a metal layer on a biaxially textured metal substrate such that the surface of the metal layer has the same texture as that of the substrate.

The present invention further provides a biaxially textured metal coated article comprising:

- 15 (a) a metal substrate; and
(b) an electrodeposited metal layer on the substrate, wherein the surface of the metal layer has the same texture as that of the substrate

The metal substrate comprises any metal which can form a suitable texture. Examples include single metals and metal alloys. The substrate is typically made by
20 rolling a metal or alloy which is generally obtained for example as a rod or sheet. Suitable metals or alloys include copper, nickel and alloys thereof, for example NiCr or NiFe. Rolling the metal or alloy forms it into a suitable article which is, for example, a tape or wire. During the rolling process, plastic flow causes reorientation of the lattice of individual grains of the substrate and the substrate tends to develop a
25 preferred orientation of the grains (the texture). The resulting article is then typically heated so as to change the crystallographic orientation of the grains. This provides a surface with a useful recrystallisation texture.

Electrodeposition is used to deposit a metal or a mixture of metals to form a textured metal surface. For example Cr, Ni, Pd, Pt, Ru, Os, Rh, Ir, Au or Cu or
30 mixtures thereof or silver may be electrodeposited by this method. Typically silver or a silver alloy is used. Silver or a silver alloy is particularly useful as an

intermediate layer between a metal substrate, for example nickel, and a superconducting layer. Suitable silver alloys include for example alloys of silver with one or more of In, Fe, Pb, Mn, Hg, Mo, Ni, Pd, Pt, Rh, Sc, Se, Au, Te, Sn, Ti, V, W, Zn, Ga, Cu, Co, Cr, Cd, As or Sb.

5 Electrodeposition can be used to deposit metals or alloys, such as those listed above, in turn to form a substrate with two or more metal layers. In one embodiment an oxygen barrier layer such as Ru, Os, Rh, Ir, Pd, Pt or Au or a mixture thereof, is deposited on the metal substrate. However, any suitable combination of metal layers may be used. Each layer may be deposited using separate baths or two layers may be
10 plated from one bath to simplify the process.

 In order to deposit more than one layer from a single plating bath, a plating bath containing all the metals to be deposited can be used. The metals should be present with compatible solution chemistries. Typically, deposition of the different metals is achieved by controlling the plating conditions, for example by controlling
15 the current density (galvanostatic control) or the cathode potential (potentiostatic control) during the plating process, generally by pulsing the potential.

 The present invention is particularly suitable for depositing a buffer layer on a substrate. The process of the present invention has the advantage that the amount of metal used to form the buffer layer can be much reduced. Further, the metal
20 develops the biaxial texture of the substrate on deposition on the substrate and it is therefore not necessary to treat the coated substrate by heating it in order to develop the texture. It is a further advantage of the present invention that lattice mismatches of the type which occur in existing processes are generally compensated for in the electrodeposition of the metal buffer layer, thereby removing the need for additional
25 buffer layers of other materials such as palladium or platinum. The nickel/silver structures of the present invention have the further advantage that the structure is thermally stable (up to 900°C). In contrast, structures using an intervening palladium layer suffer from Pd/Ag alloying during subsequent heat treatment and/or superconductor deposition.

30 The present invention is also suitable for depositing multiple metal layers on a metal substrate. In particular, one or more intermediate metal layers may be

desirable to further overcome a lattice mismatch or to act as an oxygen barrier layer. For example, when a thin layer of silver is deposited directly on a nickel substrate, oxygen can diffuse through the silver coating and react with the nickel to form nickel oxide. This can cause dewetting of the silver coating from the substrate. Depositing
5 an intermediate oxygen barrier layer, for example a layer of ruthenium, osmium, rhodium, iridium, palladium, platinum or gold, preferably rhodium, can eliminate this problem. An intermediate rhodium layer between a nickel substrate and a silver layer, for example also reduces lattice strain in the silver layer as rhodium has a lattice parameter between that of silver and nickel. Other intermediate layer metals
10 with appropriate lattice parameters can also be used to reduce lattice strain.

An oxygen barrier layer is typically deposited as an intermediate layer. For example, an oxygen barrier such as rhodium is probably not as chemically compatible with a superconductor as a metal layer such as silver. Therefore a further metal layer, such as silver, is typically deposited over the oxygen barrier layer.
15 However, in one preferred embodiment of the invention a single oxygen barrier metal layer is deposited on the metal substrate.

Thus, the present invention also provides a process which further comprises electrodepositing one or more further metal layers sequentially on the metal coated substrate, such that the or each resulting metal layer has the same surface texture as
20 that of the substrate.

The present invention also provides an article which further comprises one or more metal layers sequentially electrodeposited over the electrodeposited metal layer (b) and wherein the surface of each metal layer has the same texture as that of the substrate.

25 Overall, the process is typically reduced by several stages.

The electrodeposited layer is generally epitaxial and is coherent giving greater than 95% coverage of the substrate.

The present invention also has the advantage that the production of a suitably textured metal layer such as cube textured silver is easier and more reproducible.

30 The final composite structure obtained by the process of the present invention is also more robust than pure silver. This is important in the fabrication and operation of

any device using the final article.

Electrodeposition techniques also have the advantage that simple equipment is used and the technique is easily scalable for industrial use. In a typical embodiment the substrate is advanced through an electroplating bath at a current
5 density and speed such that the desired coating thickness is achieved. Further the technique can be performed at room temperature and atmospheric pressure which is an advantage compared to, for example, vacuum techniques.

In one aspect the process of the present invention as defined above further comprises the step of coating the, or the top, metal layer with a superconducting
10 layer. In another aspect the process includes the step of polishing or electropolishing the metal substrate before deposition of the metal layer.

The substrate is rolled using forward rolling only or using reverse rolling (the rolling direction is reversed after each pass). It is generally found that reverse rolling produces better results. The rolling speed generally influences the texture
15 development. However, its effect does not normally dominate the result achieved and higher rolling speeds are generally desirable for economic reasons. The substrate may be rolled by hand. Typically a hand-rolled substrate is rolled occasionally reversing the tape between passes. During the rolling process a lubricant is optionally employed depending on the texture required. Where a lubricant is
20 employed it is for example a light mineral oil, a heavy mineral oil, kerosene or another lubricant known for this purpose. In general, a fine grain size is desirable in the material before rolling and initial heat treatments and deformations are usually designed so as to give a random texture in the starting material before rolling.

The substrate is typically textured by annealing. Where appropriate a
25 substrate is rolled and annealed alternately in order to produce a suitable texture. A substrate is typically rolled to achieve a certain percentage deformation and then annealed. Typically a deformation of above 90% is suitable. A suitable deformation for copper is typically 95 to 97% and a suitable deformation for nickel is typically about 93%.

30 The temperature for rolling the substrate varies according to the material of the substrate and the texture that is to be produced, as one of skill in the art will be

aware. Copper rods are suitably rolled at room temperature in order to produce for example, a sharp cube texture, which is developed after annealing. Cube texture can also be achieved in alloys based on silver, copper, nickel or iron. Annealing temperatures are generally at least 50°C, for example 500 to 800°C. However, higher
5 annealing temperatures are frequently chosen. For example annealing temperatures of up to 1200°C are common. The annealing temperature will be selected for the particular metal. Thus, as an extreme example silver can recrystallise slowly at room temperature. Copper and nickel can be annealed, for example, in vacuum or in a mixture of argon and hydrogen. Copper is typically annealed at a temperature of
10 from 200-1000°C for a few seconds to an hour, for example, copper is suitably annealed in vacuum at 500 to 800°C for 1 hour. Nickel is typically annealed at a temperature of from 200 to 1200°C for a few seconds to four hours, for example nickel is suitably annealed in an argon/hydrogen mixture comprising 4% hydrogen at about 800°C for 4 hours.

15 Suitable textures for substrates include {100} <100>, {100} <110>, {110} <100> and {110} <110>. For example {100}<100> texture is typically obtained on copper or nickel substrates, in particular copper or nickel tape. The cube and hexagonal textures are often preferred, in particular the cube texture.

 The thickness of the deposited metal layer depends on the metal deposited
20 and on the proposed application for the coated substrate. Typically the metal layer has a thickness of from 1 nm to 10 µm. A silver layer electrodeposited on the substrate generally has a thickness of up to 10 µm, preferably from 0.01 to 2 µm, more preferably from 0.01 to 1 µm, for instance 0.05 to 0.5 µm, and most preferably about 0.1 to 0.2 µm. A rhodium layer generally has a thickness of from 1 nm to 10
25 µm, for instance 10 nm to 1 µm, preferably 10 nm to 250 nm.

 A metal layer of such thickness develops the texture of the substrate on deposition on the substrate. Preferably the metal layer is deposited epitaxially on the substrate. Generally the conditions of the electroplating bath are chosen so as to optimise the development of the texture of the metal layer.

30 Electrodeposition is generally carried out at a temperature of from 10 to 95°C. The current density is typically from 1 to 50000 A/m². Plating times are

generally from 10 μ s to 1 hour and the electrodeposition generally takes place at atmospheric pressure.

Electrodeposition of silver is suitably carried out at a temperature of from 10 to 95°C, more preferably 10 to 90°C, for instance 20 to 90°C, more preferably 20 to 85°C. Electrodeposition is generally conducted at atmospheric pressure. The current density used is as high as possible, generally in the range 50 to 50000 A/m², preferably 50 to 25000 A/m² and most preferably 50 to 1000 A/m². The plating time is typically from 10 μ s to 1 hour, for instance 10 seconds to 30 minutes, preferably 1 to 10 minutes, depending to an extent on the amount of current used and on the desired thickness. In addition, very fast electrodeposition processes are particularly preferred, for example with a plating time of 10 μ s to 1 s preferably 10 ms to 500 ms.

Rhodium is typically electrodeposited at a temperature of from 10 to 80°C, preferably 10 to 70°C, for instance 15 to 60°C, more preferably 20 to 50°C. The current density used is typically in the range of from 1 to 10000 A/m². The plating time is typically from 10 μ s to 1 hour, for instance 10 seconds to 30 minutes, preferably 1 to 10 minutes, depending to an extent on the current density and the desired thickness.

Once electrodeposition has taken place the texture can be studied by x-ray diffraction to confirm the texture formed.

Electrodeposition takes place in any suitable solution in order to electrodeposit the metal layer on the substrate. Such solutions will be familiar to those with knowledge of electrodeposition. A suitable plating solution typically contains a salt or oxide of the metal to be deposited, generally with a conducting salt and typically also a complex-forming agent.

A variety of silver plating solutions are known. Although solutions containing cyanide are the most common when electrodepositing silver, non-toxic solutions can also be used. Examples of silver plating solutions include hydantoin based, cyanide based, thiosulphate based and succinimide based solutions

When electrodepositing silver compounds from non-cyanide solutions the silver salt used is generally silver nitrate, silver oxide or a mixture thereof. Silver may also be used in the form of silver thiosulphate or potassium silver disuccinimide

or in the form of $\text{KAg}(\text{CN})$ in cyanide based solutions. The silver salt is generally used in combination with a conducting salt and also a complex-forming agent. The complex-forming agent may be a hydantoin compound. Suitable hydantoin compounds include 1-methylhydantoin, 1,3-dimethylhydantoin, 5,5-dimethylhydantoin, 1-methanol-5,5-dimethylhydantoin and 5,5-diphenylhydantoin. The conductive salt is typically sodium chloride, potassium chloride, potassium formate or a mixture thereof. Generally the silver is present in the bath in an amount of from 1 to 100 g/l, preferably 5 to 50g/l, more preferably 8 to 30 g/l as metal concentration. The complex-forming agent is generally present in an amount of from 10^{-15} to 10^{-2} mol/litre, preferably 10^{-10} to 10^{-3} mol/l and the conductive salt is generally present in an amount of from 1 to 100 g/l, preferably 5 to 50 g/l, more preferably 10 to 25g/l.

An example of a hydantoin based silver plating composition is silver nitrate as the silver salt, potassium chloride as the conducting salt and hydantoin as the complex-forming agent in distilled water. This composition is typically used with a pure silver anode.

An example of a cyanide based silver plating solution is silver as $\text{KAg}(\text{CN})$, potassium cyanide and potassium carbonate. This solution may be used containing for example 1 to 40 g/l $\text{KAg}(\text{CN})$, 10 to 140 g/l free potassium cyanide and 15 g/l potassium carbonate. A suitable temperature for such solutions is, for example, 20 to 30°C. The current density used for electroplating is, for example, 50 to 400 A/m^2 .

Another suitable cyanide based solution, known as a high speed silver solution, is a cyanide based solution comprising silver as $\text{KAg}(\text{CN})$ and conducting/buffering salts. Such a solution typically contains 20 to 80 g/l of silver as $\text{KAg}(\text{CN})$ and 60 to 120 g/l of conducting/buffering salts. A high speed silver solution is generally operated at a pH of from 8 to 9.5, a temperature of from 60 to 70°C and a current density of from 3000 to 38,000 A/m^2 . The solution is typically agitated rapidly while electrodeposition occurs. A suitable anode for this solution is platinum or a platinum/titanium mixture.

A suitable thiosulphate plating solution comprises, for example, silver thiosulphate, sodium thiosulphate and sodium metabisulphite. Typically such a

-10-

solution contains about 30 g/l of silver as thiosulphate, 300 to 500 g/l of sodium thiosulphate and 30 to 50 g/l of sodium metabisulphite. The solution is generally used at a pH of from 8 to 10, a temperature of from 15 to 30°C and a current density of from 40 to 100 A/m².

5 A suitable succinimide solution comprises potassium silver disuccinimide, succinimide and potassium sulphate. Typically the solution contains about 30 g/l of silver as potassium silver disuccinimide, 11 to 55 g/l of succinimide and about 45 g/l of potassium sulphate. The solution is generally used at a pH of about 8.5, a temperature of 15 to 30°C and a current density of 40 to 100 A/m².

10 Rhodium is typically electrodeposited using a rhodium sulphate solution in the presence of sulphuric acid. The solution is generally used at a pH of 2 or less preferably pH 1 or less, a temperature of 15 to 50 °C and a current density of 5 A/m². Other known rhodium plating solutions may also be used.

 The electrodeposition operates at a pH chosen according to the plating
15 chemistry used. The pH for silver deposition is, for example, from 3 to 13, preferably 5 to 11, most preferably 6 to 10. A pH of 6 to 8 is preferred for a hydantoin based solution and a pH of 9 to 10 is preferred for a cyanide based solution. A pH of 2 or less is generally used for rhodium.

 A particularly preferred procedure is to electrodeposit a silver layer on a
20 nickel or nickel alloy substrate. Other preferred procedures include depositing a rhodium layer, or a rhodium layer and then a silver layer, on a copper or nickel substrate.

 The metal substrate is generally electropolished before the metal layer is electrodeposited on the substrate. However, electropolishing is not always necessary.
25 The use of an intermediate metal layer may remove the need for electropolishing the substrate. In particular, a very acidic plating bath may remove any residual oxide from the surface of the textured metal substrate. Any conventional electropolishing technique known in the art may be used. For example a solution of phosphoric acid or sulphuric acid may be appropriate. Typical electropolishing times range from 5
30 seconds to 10 minutes at a current density of about 5000 A/m². A suitable electropolishing technique will be chosen having regard to the metal substrate being

-11-

used; suitable techniques will be known by those skilled in the art. For example copper frequently does not need to be electropolished. However, if desired, copper can be electropolished using phosphoric acid, for example 70% phosphoric acid for from 5 to 60 seconds at a current density of 5000 A/m². Nickel and nickel alloys are
5 generally electropolished using sulphuric acid. Typically nickel is electropolished using a solution of 4H₂SO₄.3H₂O for about 3.5 minutes at 6000 A/m². A Ni-10Cr alloy is typically suitably electropolished for 10 minutes at 6000 A/m² in 4H₂SO₄.3H₂O. A nickel-30Fe alloy is typically suitably electropolished in 4H₂SO₄.3H₂O for 5 minutes at 6000 A/m².

10 The coated substrate can then be coated with a superconducting layer. Suitable superconductors include superconductors from the Re-Ba-Cu-O (Re denotes a rare earth element), Tl-(Pb, Bi)-Sr-(Ba)-Ca-Cu-O and Hg-(Pb)-Sr-(Ba)-Ca-Cu-O families. A typical superconductor for deposition on the metal layer is YBaCuO. NdBCO and Tl,Bi 1223 are also commonly used. MgB₂ may also be used. The
15 superconductor is deposited on the metal layer by any suitable method known in the art, for example sputtering, pulsed laser deposition, BaF₂ precursor coevaporation, e-beam evaporation, MOCVD, liquid phase epitaxy, spray pyrolysis, sol-gel or electrodeposition.

In one embodiment the metal coated substrate is coated with a ceramic buffer
20 layer or superconducting layer and then one or more further metal layers are electrodeposited on top. The ceramic layer may be a single crystal or a textured layer. Thus a multilayer structure comprising metal and ceramic or superconducting layers is built up. The final deposited layer is a metal layer. In a particular embodiment the substrate is coated in order by one or more metal layers, a ceramic
25 layer, one or more metal layers, a ceramic buffer layer or superconducting layer and a final metal layer.

The invention is further illustrated by the following examples:

Examples

30 Example 1.

A 99.98% pure sample of oxygen-free high-conductivity copper with a

-12-

thickness of 2 mm was obtained. The copper was rolled to achieve 95% deformation and then annealed at 700°C for 1 hour in order for texture to develop. The copper substrate was then polished using a 70% phosphoric acid solution in water and a current density of 5000 A/m² for 1 minute.

- 5 A metal layer of silver was electrodeposited on the copper substrate using a plating bath with the following composition:

AgNO₃ 8 g/l

KCl 16 g/l

- 10 hydantoin 40 g/l

which was made up to concentration using distilled water. The electrodeposition took place at 57°C using a current density of 500 A/m² for 15 minutes. The anode was made of pure silver. The pH of the bath started at 6 and rose to pH 7 during the electrodeposition process.

- 15 The texture of the deposited silver layer was confirmed by x-ray diffraction.

Example 2.

- A sample of pure nickel was rolled to achieve 93% deformation and then annealed at 800°C for 4 hours to allow the texture to develop. The nickel substrate
20 was electropolished using 4H₂SO₄.3H₂O in an electropolishing bath for 3.5 minutes at a current density of 6000 A/m².

A metal layer of silver was then deposited on the nickel substrate using the plating bath composition of Example 1.

- Electrodeposition was carried out for 15 minutes at a current density of 500
25 A/m² and a temperature of 53°C. The pH of the bath was pH 7 throughout the deposition practice. The anode was pure silver.

The texture of the metal layer was confirmed by x-ray diffraction.

Example 3.

- 30 A 99.98% pure sample of oxygen-free high-conductivity copper was obtained with a thickness of 2mm. The copper sample was rolled to achieve 95%

deformation. The sample was then annealed at 500°C for 1 hour in order for the texture to develop.

A metal layer of silver was then electrodeposited on the copper sample without electropolishing the copper sample. The electrodeposition bath had the same composition as in Example 1.

The electrodeposition process was carried out for 30 minutes at a temperature of 80°C. The pH of the bath started at pH 6 and rose to pH 7 during the deposition process. The anode was pure silver.

The texture of the metal layer was confirmed by x-ray diffraction.

Example 4.

A 90% nickel-10% chromium alloy sample was rolled to achieve 97.5% deformation. The sample was then annealed to 1000°C for 4 hours to allow the texture to form. The nickel/chromium alloy was electropolished using 4H₂SO₄·3H₂O electropolishing bath. Electropolishing was carried out for 10 minutes at a current density of 6000 A/m².

The nickel-chromium alloy was electrodeposited with a silver layer using an electrodeposition bath with a composition as in Example 1.

Electrodeposition was carried out at 65°C for 10 minutes at a current density of 500 A/m². The pH of the electrodeposition bath remained at pH 7 throughout the deposition process. The anode was made of pure silver.

The texture of the silver layer was confirmed by x-ray diffraction.

Example 5.

A 99.98% pure sample of oxygen-free high-conductivity copper with a thickness of 2 mm was obtained. The copper was rolled to achieve 97% deformation and then annealed at 800°C for 1 hour in order for texture to develop. The copper substrate was then polished using a 70% phosphoric acid solution in water and a current density of 20000 A/m² for 20 seconds.

A metal layer of rhodium was electrodeposited on the copper substrate using a plating bath with the following composition:

-14-

Rhodium (as $\text{Rh}_2(\text{SO}_4)_3 \cdot 12\text{H}_2\text{O}$) 2 g/l
 H_2SO_4 20 m/l

the solution was made up using Johnson Matthey RH8 rhodium sulphate plating solution concentrate (8g Rh/100 ml) and de-ionised water. The electrodeposition
5 took place at a current density of 50 A/m² for 50 seconds. The anode was made of platinized titanium.

The texture of the deposited rhodium layer was confirmed by x-ray diffraction.

10 Example 6.

A 99.99% pure nickel rod of 5 mm diameter was rolled to achieve 93% deformation and then annealed at 800°C in a mixture of argon and 4% hydrogen for 4 hours to allow the texture to develop. The nickel substrate was electropolished using $4\text{H}_2\text{SO}_4 \cdot 3\text{H}_2\text{O}$ in an electropolishing bath for 4 minutes at a current density of
15 6000 A/m².

A metal layer of rhodium was then deposited on the nickel substrate using the plating bath composition of Example 5.

Electrodeposition was carried out for 8 minutes at a current density of 5 A/m². The anode was platinized titanium.

20 The texture of the metal layer was confirmed by x-ray diffraction.

Example 7.

A 99.99% pure nickel rod of 5 mm diameter was rolled to achieve 93% deformation and then annealed at 800°C in a mixture of argon and 4% hydrogen for
25 4 hours to allow the texture to develop. The nickel substrate was electropolished using $4\text{H}_2\text{SO}_4 \cdot 3\text{H}_2\text{O}$ in an electropolishing bath for 4 minutes at a current density of 6000 A/m².

A metal layer of rhodium was then deposited on the nickel substrate using the plating bath composition of Example 5. The electrodeposition took place at 20°C
30 and pH 1.

Electrodeposition was carried out for 5 minutes at a current density of 5 A/m².

-15-

The anode was platinized titanium.

A metal layer of silver was then deposited on the rhodium layer using the plating bath composition of Example 1.

The electrodeposition took place at 20°C for 30 seconds using a current
5 density of 500 A/m² at pH 10. The anode was pure silver.

The texture of the silver layer was confirmed by x-ray diffraction.

Example 8.

A 99.98% pure sample of oxygen-free high-conductivity copper with a
10 thickness of 2 mm was obtained. The copper was rolled to achieve 97% deformation and then annealed at 800°C for 1 hour in order for texture to develop. The copper substrate was then polished using a 70% phosphoric acid solution in water and a current density of 20000 A/m² for 20 seconds.

A metal layer of rhodium was electrodeposited on the copper substrate using
15 a plating bath with the following composition:

Rh ₂ (SO ₄) ₃ ·12H ₂ O	2 g/l
H ₂ SO ₄	20 m/l

the solution was made up using Johnson Matthey RH8 rhodium sulphate plating solution concentrate (8g Rh/100 ml) and de-ionised water. The electrodeposition
20 took place at a current density of 50 A/m² for 1 minute at 20°C and pH 1. The anode was made of platinized titanium.

A metal layer of silver was electrodeposited on the rhodium layer using a plating bath with the following composition:

Ag ₂ SO ₄	30 g/l
25 25% NH ₄ OH	75 g/l
KI	600 g/l
Na ₄ P ₂ O ₇	60 g/l

which was made up to concentration using distilled water. The electrodeposition took place using a current density of 100 A/m² for 30 seconds at 20°C and pH 10.
30 The anode was made of pure silver.

The texture was confirmed by x-ray diffraction.

CLAIMS

1. A process for producing a metal article coated with a metal layer having a biaxially textured surface, which process comprises electrodepositing the
5 metal layer on a biaxially textured metal substrate such that the surface of the metal layer has the same texture as that of the substrate.
2. A process according to claim 1 wherein the metal layer is an oxygen barrier layer.
10
3. A process according to claim 2 wherein the oxygen barrier layer is chosen from rhodium, osmium, ruthenium, iridium, palladium, platinum and gold.
4. A process according to any one of the preceding claims which further
15 comprises electrodepositing one or more further metal layers sequentially on the metal coated metal article, such that the or each resulting metal layer has the same surface texture as that of the substrate.
5. A process according to claim 1 or 4 wherein the, or the further, metal
20 layer is silver.
6. A process according to any one of the preceding claims wherein no heat treatment is used to obtain the textured surface on the or each metal layer.
- 25 7. A process according to any one of the preceding claims wherein the or each metal layer has a thickness of up to 10 μm .
8. A process according to claim 7 wherein the or each metal layer has a
thickness of from 0.01 to 1 μm .
30
9. A process according to any one of the preceding claims wherein the

-17-

substrate is nickel or a nickel alloy.

10. A process according to any one of the preceding claims wherein the or each electrodeposition takes place at a temperature from 10 to 90°C.

5

11. A process according to any one of the preceding claims wherein the silver is electrodeposited using a current density of from 50 to 50,000 A/m².

12. A process according to any one of claims 1 to 10 wherein rhodium is electrodeposited using a current density of from 1 to 10000 A/m².

10

13. A process according to any one of claims 1 to 11 wherein silver is electrodeposited at a pH of from 6 to 10.

15

14. A process according to any one of claims 1 to 10 or 12 wherein rhodium electrodeposited at a pH of 2 or less.

15. A process according to any one of the preceding claims wherein the or each electrodeposition takes place for from 10 µs to 1 hour.

20

16. A process according to any one of the preceding claims wherein the metal substrate is electropolished before the metal layer is electrodeposited on it.

17. A process according to any one of the preceding claims which further comprises coating the metal coated metal article with a superconducting or ceramic layer.

25

18. A process according to claim 17 which further comprises electrodepositing a metal layer on the superconducting or ceramic layer.

30

19. A biaxially textured metal coated article comprising:

-18-

- (a) a metal substrate; and
- (b) an electrodeposited metal layer on the substrate, wherein the surface of the metal layer has the same texture as that of the substrate.

5 20. An article according to claim 19 which further comprises one or more metal layers sequentially electrodeposited over the electrodeposited metal layer (b) and wherein the surface of each metal layer has the same texture as that of the substrate.

INTERNATIONAL SEARCH REPORT

International Application No
PCT/GB 01/01793

A. CLASSIFICATION OF SUBJECT MATTER

IPC 7 C25D7/00 C25D5/10 H01L39/24

According to International Patent Classification (IPC) or to both national classification and IPC

B. FIELDS SEARCHED

Minimum documentation searched (classification system followed by classification symbols)

IPC 7 C25D H01L

Documentation searched other than minimum documentation to the extent that such documents are included in the fields searched

Electronic data base consulted during the international search (name of data base and, where practical, search terms used)

INSPEC, COMPENDEX, EPO-Internal, WPI Data, PAJ

C. DOCUMENTS CONSIDERED TO BE RELEVANT

Category *	Citation of document, with indication, where appropriate, of the relevant passages	Relevant to claim No.
X	GB 1 328 016 A (PHILIPS ELECTRONIC ASSOCIATED) 22 August 1973 (1973-08-22) page 1, line 51 -page 2, line 62 page 2; table --- -/--	1-4,6-8, 10,11, 19,20

☒ Further documents are listed in the continuation of box C.

☒ Patent family members are listed in annex.

* Special categories of cited documents:

- *A* document defining the general state of the art which is not considered to be of particular relevance
- *E* earlier document but published on or after the international filing date
- *L* document which may throw doubts on priority claim(s) or which is cited to establish the publication date of another citation or other special reason (as specified)
- *O* document referring to an oral disclosure, use, exhibition or other means
- *P* document published prior to the international filing date but later than the priority date claimed

- *T* later document published after the international filing date or priority date and not in conflict with the application but cited to understand the principle or theory underlying the invention
- *X* document of particular relevance; the claimed invention cannot be considered novel or cannot be considered to involve an inventive step when the document is taken alone
- *Y* document of particular relevance; the claimed invention cannot be considered to involve an inventive step when the document is combined with one or more other such documents, such combination being obvious to a person skilled in the art.
- *G* document member of the same patent family

Date of the actual completion of the international search

13 August 2001

Date of mailing of the international search report

27/08/2001

Name and mailing address of the ISA
European Patent Office, P.B. 5818 Patentlaan 2
NL - 2280 HV Rijswijk
Tel. (+31-70) 340-2040. Tx. 31 651 epo nl.
Fax: (+31-70) 340-3016

Authorized officer

Zech, N

INTERNATIONAL SEARCH REPORT

International Application No

PCT/GB 01/01793

C.(Continuation) DOCUMENTS CONSIDERED TO BE RELEVANT		
Category *	Citation of document, with indication, where appropriate, of the relevant passages	Relevant to claim No.
P,X	<p>DATABASE INSPEC 'Online! INSTITUTE OF ELECTRICAL ENGINEERS, STEVENAGE, GB; GOODALL R ET AL: "Fabrication of cube-textured Ag-buffered Ni substrates by electro-epitaxial deposition" Database accession no. 6948084 XP002174798 abstract & SUPERCONDUCTOR SCIENCE & TECHNOLOGY, MARCH 2001, IOP PUBLISHING, UK, vol. 14, no. 3, pages 124-129, ISSN: 0953-2048</p> <p>---</p>	1,2,5-9, 19
X	<p>PATENT ABSTRACTS OF JAPAN vol. 1999, no. 08, 30 June 1999 (1999-06-30) & JP 11 060398 A (RES DEV CORP OF JAPAN), 2 March 1999 (1999-03-02)</p> <p>---</p>	1,5,6,19
A	<p>abstract</p> <p>---</p>	2,3
X	<p>DATABASE COMPENDEX 'Online! ENGINEERING INFORMATION, INC., NEW YORK, NY, US; FISHER JE: "STRUCTURE AND MAGNETIC ANISOTROPY OF F.C.C. COBALT FILMS BETWEEN 150 AND 870 K" Database accession no. EIX71020023638 XP002174799 abstract & THIN SOLID FILMS JAN 1970, vol. 5, no. 1, January 1970 (1970-01), pages 53-60,</p> <p>---</p>	1,6,19
X	<p>H.W. DETTNER AND J. ELZE: "Handbuch der Galvanotechnik" 1963, CARL HANSER VERLAG, MÜNCHEN XP002173627 page 93, paragraph 3 -page 95, paragraph 1</p> <p>---</p>	1,19
A	<p>US 5 739 086 A (BUDAI JOHN D ET AL) 14 April 1998 (1998-04-14)</p> <p>abstract column 4, line 1-7,44-54 column 10, line 55-60 example 10</p> <p>---</p> <p style="text-align: center;">-/--</p>	1-3,5,7, 9,17,19, 20

INTERNATIONAL SEARCH REPORT

International Application No
PCT/GB 01/01793

C.(Continuation) DOCUMENTS CONSIDERED TO BE RELEVANT		
Category *	Citation of document, with indication, where appropriate, of the relevant passages	Relevant to claim No.
A	<p>PATENT ABSTRACTS OF JAPAN vol. 1996, no. 02, 29 February 1996 (1996-02-29) & JP 07 278881 A (NIPPON STEEL CORP), 24 October 1995 (1995-10-24) abstract -& DATABASE WPI Derwent Publications Ltd., London, GB; AN 1995-401552 XP002173628 abstract</p> <p>-----</p>	1,19

INTERNATIONAL SEARCH REPORT

Information on patent family members

International Application No

PCT/GB 01/01793

Patent document cited in search report	Publication date	Patent family member(s)	Publication date
GB 1328016 A	22-08-1973	NL 7009689 A CA 924194 A DE 2131231 A FR 2098125 A	04-01-1972 10-04-1973 13-01-1972 03-03-1972
JP 11060398 A	02-03-1999	JP 2980869 B	22-11-1999
US 5739086 A	14-04-1998	US 5741377 A AU 713892 B AU 5539896 A CA 2217822 A EP 0830218 A JP 11504612 T US 5958599 A WO 9632201 A US 6235402 B US 5898020 A US 5968877 A	21-04-1998 16-12-1999 30-10-1996 17-10-1996 25-03-1998 27-04-1999 28-09-1999 17-10-1996 22-05-2001 27-04-1999 19-10-1999
JP 07278881 A	24-10-1995	NONE	

**This Page is Inserted by IFW Indexing and Scanning
Operations and is not part of the Official Record**

BEST AVAILABLE IMAGES

Defective images within this document are accurate representations of the original documents submitted by the applicant.

Defects in the images include but are not limited to the items checked:

- ☐ **BLACK BORDERS**
- ☐ **IMAGE CUT OFF AT TOP, BOTTOM OR SIDES**
- ☐ **FADED TEXT OR DRAWING**
- ☐ **BLURRED OR ILLEGIBLE TEXT OR DRAWING**
- ☐ **SKEWED/SLANTED IMAGES**
- ☐ **COLOR OR BLACK AND WHITE PHOTOGRAPHS**
- ☐ **GRAY SCALE DOCUMENTS**
- ☐ **LINES OR MARKS ON ORIGINAL DOCUMENT**
- ☐ **REFERENCE(S) OR EXHIBIT(S) SUBMITTED ARE POOR QUALITY**
- ☐ **OTHER:** _____

IMAGES ARE BEST AVAILABLE COPY.

As rescanning these documents will not correct the image problems checked, please do not report these problems to the IFW Image Problem Mailbox.